

The 2nd Law of Thermodynamics Refutes Evolution!

Part 1

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Introduction

There have been a few discussions that I have run across concerning the Second Law of thermodynamics and closed systems. I feel there is a misunderstanding of the concept of open system verses a closed system, and misconception that the Second Law only applies to closed systems. I want to describe here why it is necessary to use a closed system and how this does prove that we are in an ever increasing state of entropy (chaos and disorder) contrary of what evolutionist promoters would have you believe.

Lets first start by examining the Laws of thermodynamics:

Zeroth Law of Thermo Dynamics

Thermal equilibrium – The most reliable definition of temperature is that it is a numerical scale for uniquely ordering the ?hotness? of a series of objects. A universal scale of temperature can be developed because of the Zeroth Law of thermodynamics: if two objects are in thermo equilibrium with a third, then they are in equilibrium with each other. The Zeroth law is a law in the sense that it is a fact of experience that must be regarded as an empirical fact of nature.

1st Law Thermo Dynamics

The total energy of the system plus the surroundings will not change and energy is conserved. (Conservation of energy) Enthalpy - deals with the heat absorbed or released under constant pressure. Generally it is favorable for the enthalpy of a system to decrease by transferring the heat to the surroundings.

2nd Law Thermodynamics

There is an inherent direction in which processes occur. A process proceeds in a direction such that the unavailable energy (the entropy) increases. In other words, in any closed system, the amount of disorder always increases with time. Things progress naturally from order to disorder, etc. Entropy (S) - is related to randomness or disorder; the greater the entropy the greater the disorder. Chemical systems show that the same behavior, namely the entropy of the universe must increase during spontaneous processes even if the entropy of the system decreases.

3rd Law Thermodynamics

If we decrease the thermal energy of a system by lowering the temperature, the energy stored in transitional, vibrational, and rotational forms of motion decrease as less energy is stored, the entropy of the system decreases. If we keep lowering the temperature do we reach a state in which these motions are essentially shut down, a point of perfect order.

The 3rd Law of thermodynamics addresses this question, which states that the entropy of a pure crystalline substance at absolute zero is zero $s(0K)=0$.

Discussion

There are fundamental laws in the Universe in which God has established to which there is no known exception, three of which are the Laws of Thermodynamics. When there is any work done due to energy conservation, there is always some dissipation of useful energy. This means, according to thermodynamics, for a closed system the measure of energy no longer available for useful work is increasing, this is called entropy. Thus the overall entropy of a closed system is increasing. This law applies to all areas of science, mechanics, engines, and even living systems. Cars rust, machines wear out, organic compounds break down, living systems die, etc. No spontaneous reversal of this process has ever been observed. An example of living systems is that which is dead (such as a leaf or stick from a tree) has no information or teleonomy¹ within it to convert the sun's energy to useful work. The heat from the sun's energy will just heat up the dead leaf or stick and entropy will increase (decay). Now, let's look at why it is necessary to create a closed system. It is absolutely necessary to create a closed system so that we can observe the Second Law of thermodynamics in action (A system protected from any outside confusing inputs). This way it is possible to see what is happening in the system that is independent of outside events. Now what happens in an open system is that at any point we see the sum of all the different downward trends acting at once. Some people have proposed that the second law only applies to closed systems, this is a classic example of their lack of understanding and confusion of the experimental necessity for a closed system to test for the existence of the Second Law, with the actual actions of the Second Law being evident in the open systems in which we live. Because the universe itself may be considered to be an isolated system and all naturally occurring processes are irreversible, one statement of the Second Law of Thermodynamics says that the entropy of the universe is constantly increasing. In fact, Clausius summarized the first two laws for thermodynamics by: "The energy of the Universe is constant; the entropy is tending to maximum."²

1. Irreversible processes will result in an increase in entropy of the universe. (Irreversible process will result in entropy generation.) Irreversible processes result in loss capability for performing work.³
2. Reversible processes result in no increase in entropy of the universe. (Reversible processes result in no entropy generation.)³
3. Proposed processes which would result in a decrease of entropy of the universe are impossible. (Impossible processes result in negative entropy generation.)³ In conclusion, we along with this entire universe are on a downward spiral towards an ever-increasing state of entropy (chaos). The Second Law of thermodynamics proves this point. And closed systems are an experimental necessity for us to test

for and to observe the existence of the Second Law of thermodynamics that is evident in the universe we live in.

EVOLUTION IS IMPOSSIBILITY!

References & Notes

1. Teleonomy, n. concept that a body or organism's structure has intrinsic value (i.e. enabling the organism to survive)
2. Donald A. McQuarrie, John D. Simon, *Physical Chemistry A Molecular Approach* (University Science Books, 1997), p. 829.
3. J. Richard Elliot, Carl T. Lira, *Introduction Chemical Engineering Thermodynamics* (Prentice Hall PTR, 1999), p. 97.